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研究課題名（和文）Development of separation method for secondary copper and arsenic minerals using flotation

研究課題名（英文）Development of separation method for secondary copper and arsenic minerals using flotation

研究代表者

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研究成果の概要（和文）：本研究は、砒素を含む鉱物から銅を分離するプロセスを、浮遊法と過酸化水素を用いた酸化処理によって実証することに成功した。研究は、小規模な実験室実験から大規模なベンチ実験に至るまで、銅-砒素精鉱に対する酸化処理の基礎的な調査および実用的な実施を含んでいた。この研究により、最終銅精鉱中のヒ素濃度を、ヒ素を含む銅精鉱の処理基準値（0.2%）より低い値まで低減できることが実証された。さらに本研究は、ヒ素を含む銅精鉱から精製銅精鉱を生成する方法を提案している。

研究成果の学術的意義や社会的意義

This research developed a process to separate arsenic from copper concentrate using flotation and oxidation treatment, producing a clean copper concentrate that contains a minimal amount of arsenic. This will decrease arsenic emissions and minimize the pollution of the environment with arsenic.

研究成果の概要（英文）：This study successfully demonstrated the process of separating copper from minerals containing arsenic via flotation and oxidation treatment using hydrogen peroxide. The research included the basic investigation and practical implementation of the oxidation treatment on copper-arsenic concentrate, ranging from small-scale laboratory experiments to larger-scale bench experiments. This research demonstrated the ability to decrease the arsenic level in the final copper concentrate to a value lower than the threshold for processing copper concentrate with arsenic, which is set at 0.2%. Moreover, this research proposes a method for generating a purified copper concentrate from copper concentrate that contains arsenic.

研究分野：Mineral processing and flotation

キーワード：Flotation Enargite Chalcocite Copper arsenic Oxidation treatment Hydrogen peroxide

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1. 研究開始当初の背景

Arsenic is an undesirable element in copper concentrate because of its toxicity and detrimental effects on human health and the environment, which diminishes the quality and value of the copper concentrate. Prior research conducted by Suyantara et al. (2020, Minerals Engineering, 152) has demonstrated that the application of oxidation treatment in selective flotation is a highly promising method for effectively separating arsenic-bearing copper minerals (enargite and tennantite) from primary copper sulfide minerals (chalcopyrite and bornite). Nevertheless, there is a scarcity of research on the specific separation of copper minerals containing arsenic and secondary copper sulfide minerals, such as chalcocite. Furthermore, there is a lack of detailed reports on the implementation of oxidation treatment for a copper-arsenic concentrate with higher complexity.

2. 研究の目的

The aim of this research is to examine the effect of oxidation treatment using hydrogen peroxide (H_2O_2) on the floatability of chalcocite and enargite, both in the absence and presence of potassium amyl xanthate (KAX). This research examined the use of H_2O_2 treatment in the process of selectively separating chalcocite and enargite using flotation. In addition, this research investigates the impact of H_2O_2 treatment on the separation of arsenic and copper sulfide minerals in a complex copper-arsenic concentrate. This research aimed to identify the optimal conditions for producing a copper concentrate with minimal arsenic contamination and the greatest copper grade.

3. 研究の方法

Pure chalcocite and enargite were crushed, hand-picked, dry-ground, and dry-sieved to obtain a particle size of 38–74 μm . Both chalcocite and enargite were mixed and then treated with KAX and various concentrations of H_2O_2 at pH 9 for 1 hour. The flotation test was carried out using a Partridge-Smith flotation glass cell and nitrogen gas. The froth was collected for 6 minutes, and the recovery was calculated based on the weight fraction of the dried froth.

A complex copper-arsenic concentrate was prepared with a pulp density of 37% (w/w). The suspension was agitated in Fahrenwald flotation equipment for 1 hour. Subsequently, the oxidation treatment was performed using 0.1 M H_2O_2 for 1 h. KAX was added, and the flotation test was performed. The froth concentrate was collected for 30 minutes and dried. The copper and arsenic recoveries were calculated and reported in this study.

4. 研究成果

Figure 1 demonstrates the effect of H_2O_2 concentration on the floatability of chalcocite and enargite. The recovery of chalcocite had a modest increase after the oxidation treatment using a 1 mM H_2O_2 aqueous solution. However, the recovery of enargite had a substantial rise from 17% to 77% after the oxidation treatment using a 1 mM aqueous solution. The findings suggest that using a lower concentration of H_2O_2 during oxidation treatment may increase the capacity of enargite to float and preferentially separate it from chalcocite. Nevertheless, the use of intense oxidation treatment might hinder the capacity of both minerals to float and decrease the degree of separation selectivity. Figure 1 shows that the separation of chalcocite and enargite was most effective, reaching 50%, following the oxidation treatment using a 1 mM H_2O_2 solution at pH 9.

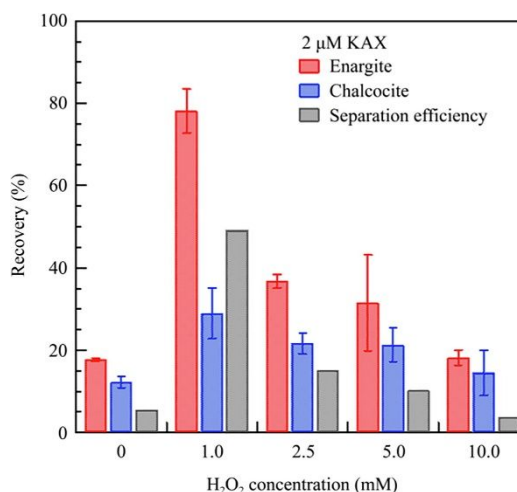


Figure 1. Effect of H_2O_2 treatment on the recovery of mixed chalcocite and enargite.

Figure 2 shows the effect of H_2O_2 oxidation treatment on the floatability of copper and arsenic sulfide minerals in a complex copper-arsenic concentrate. Figure 2 shows that

the recoveries of total copper, arsenic, and non-arsenic copper after treatment with 60 g/t of KAX increased with increasing pH. The recovery of arsenic significantly improved by 29% after increasing the pH value from 9 to 10. The recovery of arsenic was 99% after treatment with 60 g/t of KAX at pH 10. The recovery of non-arsenic copper after increasing the pH value from 9 to 10 increased by 17%, which is significantly lower than the improvement in arsenic recovery. This is likely due to the more rapid adsorption of KAX onto the surface of arsenic-bearing copper minerals compared to that onto the surface of copper sulfide minerals at a higher pH value (alkaline condition).

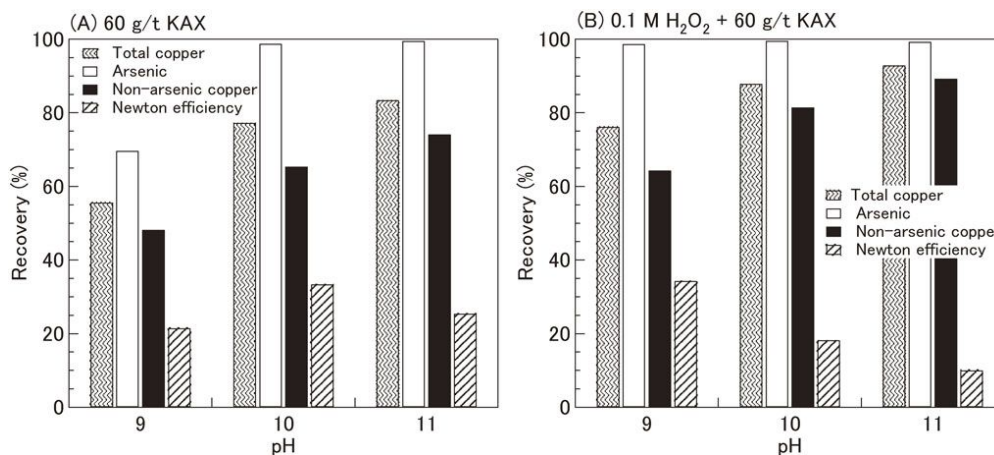


Figure 2. Effect of H₂O₂ oxidation treatment on the floatability of copper and arsenic sulfide minerals in a complex copper-arsenic concentrate at various pH conditions.

As shown in Figure 2, the optimum conditions (33%) for separating copper sulfide and arsenic-bearing copper minerals using 60 g/t of KAX was at pH 10. The Newton efficiency decreased at pH 11. This is due to a higher recovery of non-arsenic copper and the maximum recovery of arsenic. The arsenic level sharply decreased from 2.32% at pH 9 to 0.14% at pH 10 and 0.06% at pH 11 after treatment with 60 g/t of KAX and 30 min of flotation. These results indicate that producing a copper concentrate with a lower grade of arsenic, that is, lower than the penalty limit of 0.2%, is possible using the KAX treatment at the pH values of 10 and 11. These results demonstrate that KAX treatment under higher pH conditions can produce a “clean” copper concentrate with a lower arsenic grade. However, it should be noted that the grade of copper also decreased with increasing pH, that is, from 25% at pH 9 to 17% at pH 10 and 15% at pH 11.

Figure 2(B) presents the effect of oxidation treatment using a 0.1 M H₂O₂ aqueous solution and adding 60 g/t of KAX under various pH conditions on the recoveries of total copper, arsenic, and non-arsenic copper and Newton efficiency. The recovery of arsenic reached its maximum after 30 min of flotation at pH 9. Therefore, increasing the pH had no significant effect on the recovery of arsenic. In contrast, the recovery of non-arsenic copper increased with increasing pH. Accordingly, the Newton efficiency decreased with increasing pH. For instance, the Newton efficiencies were 18% and 10% at pH values of 10 and 11, respectively. Furthermore, the grade of copper decreased from 20% at pH 9 to 15% at pH 10 and 10% at pH 11 (Fig. 10(A)). Meanwhile, the grade of arsenic decreased slightly from 0.16% at pH 9 to 0.07% and 0.15% at pH values of 10 and 11, respectively. These results indicate that oxidation treatment using H₂O₂, followed by the addition of KAX at a higher pH was less effective for separating copper and arsenic-bearing copper minerals.

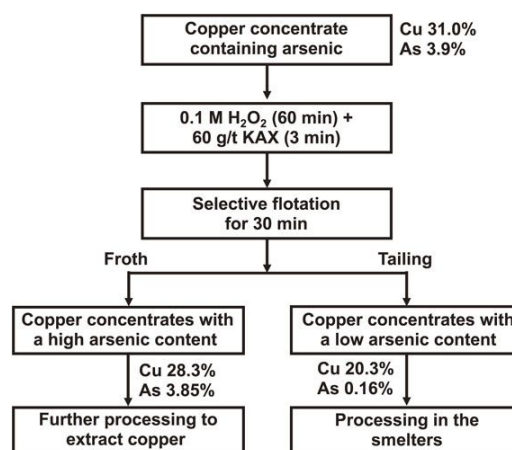


Figure 3. Proposed selective flotation process for the separation of arsenic and copper concentrate.

Based on these flotation results, we proposed a new process to separate copper and arsenic-bearing copper minerals via flotation, as shown in Figure 3. Selective flotation produces a “dirty” copper concentrate, which contains copper with high-grade arsenic, and a “clean” copper concentrate, which contains copper with low-grade arsenic in the tailing. The “dirty” copper concentrate can be further processed to extract copper from arsenic-bearing copper minerals, whereas the “clean” copper concentrate can be processed in smelters.

5. 主な発表論文等

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〔図書〕 計0件

〔産業財産権〕

〔その他〕

資源処理・環境修復工学研究室 https://process.mine.kyushu-u.ac.jp/

6. 研究組織		
氏名 (ローマ字氏名) (研究者番号)	所属研究機関・部局・職 (機関番号)	備考

7. 科研費を使用して開催した国際研究集会

〔国際研究集会〕 計0件

8. 本研究に関連して実施した国際共同研究の実施状況

共同研究相手国	相手方研究機関
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